

ELECTROCHEMICAL GENERATOR HAVING AN ALL-SOLID-STATE  
POLYMERIC ELECTROLYTE COMPRISING FLUOROPOLYMERS

The invention relates to the field of rechargeable  
5 lithium storage batteries or all-solid-state electro-  
chemical generators, of the type comprising at least  
one negative electrode capable of delivering a lithium  
cation, an all-solid-state alkaline polymeric  
electrolyte and a positive electrode capable of  
10 incorporating the nonionized species corresponding to  
said lithium cation.

The invention also relates to the all-solid-state  
polymeric electrolytes useful, in particular, for  
15 producing the electrochemical generators according to  
the invention.

The operation of a lithium storage battery involves the  
transfer by ionic conduction, via a plastic or liquid  
20 electrolyte, of lithium cations coming from the  
negative electrode or "source" to the positive  
electrode or "well" in the case of the nonionized  
species corresponding to the lithium cation.

25 In the case of rechargeable storage batteries called  
secondary storage batteries, it is known that these  
have to have an almost constant specific energy during  
the many charge/discharge cycles.

30 In practice, a storage battery must be able to undergo  
more than 500 charge/discharge cycles without the  
delivered energy being significantly reduced.

A problem likely to affect the constancy of the energy  
35 delivered during charge/discharge cycles lies in the  
imperfect deposition of lithium on the lithium negative  
electrode. It has in fact been found that in lithium  
storage batteries the deposition of lithium during  
recharging occurs nonuniformly, in the form of tree

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structures or dendrites, which gives rise to local short circuits. It is acknowledged that this phenomenon occurs more rapidly the higher the current density. This phenomenon limits the lifetime, that is to say the  
5 number of charge/discharge cycles, of storage batteries.

The use of a polymeric electrolyte partly overcomes this problem.

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Two technologies are used at the present time:

- all-solid-state or "dry" technology;
- plasticized or gelled technology.

15 The addition of a plasticizer is justified by the substantial improvement in the ionic conductivity of the electrolytic membrane. Operation at room temperature, or even lower, becomes possible. This is far from being the case with all-solid-state  
20 technology.

The addition of a plasticizer requires the incorporation of another polymer. This is because the mechanical strength of polyethers (frequently used in  
25 both these technologies) is too low to allow use as a separator when a plasticizer is incorporated into them. This polymer is in general a fluoropolymer. A ratio of 1 between the polyether and the fluoropolymer is a good compromise between conductivity and mechanical strength  
30 (see US 6 185 645). The incorporation of too large an amount of fluoropolymer will have deleterious consequences on the conductivity, since a fluoropolymer is much inferior in terms of ionic conductivity to polyethers.

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In the case of all-solid-state technology, the mechanical strength is provided by the polyether itself. Its mechanical strength is sufficient and does not require the incorporation of another polymer. The

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incorporation of a fluoropolymer even becomes deleterious from the standpoint of ionic conductivity.

In general, such lithium storage batteries result from the lamination/assembly of three thin films (three-layer assembly): a film of positive electrode containing an electrochemically active material, a film of alkaline polymeric electrolyte, especially a polyether, and of a lithium salt, and a film of a lithium-based negative electrode.

The storage battery is connected up by a collector associated with the positive electrode, the negative electrode itself acting as collector.

The thickness of such a storage battery is around 30 to 300  $\mu\text{m}$ , each of the electrode films having a thickness of 10 to 100  $\mu\text{m}$ . It should be noted that, the polymeric electrolyte essentially acting as a cation transporter, its thickness may be thin, especially much thinner than the electrodes with which it is associated.

To further limit the formation of dendrites, it has been proposed to modify the surface of the lithium anode by hydrofluoric acid (Takehara: 8th International Congress, Nagoya 1996). This treatment of the lithium anode substantially improves the performance of the cells - fluorine modifies the oxidized surface layer of the lithium, thereby reducing the reactivity of the lithium with respect to the electrolyte.

It has also been proposed to incorporate  $\text{CO}_2$  (Z. Takehara et al., *J. Power Sources*, 43/44, 377 (1993)).

It is by a completely different route that the inventors have solved the abovementioned problem.

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It is one of the objects of the present invention to propose novel all-solid-state polymeric electrolytes making it possible to ensure very many charge/discharge cycles at a virtually constant specific energy, in particular by reducing the tree-forming phenomenon during lithium redeposition on the negative lithium electrode.

Moreover, these novel all-solid-state polymeric electrolytes are easy to manufacture and possess excellent mechanical properties.

The invention is based on the observation that adding, to the all-solid-state polymeric electrolytes, in addition to the possible usual fillers, small amounts of fluoropolymers allows the desired results mentioned above to be achieved.

The invention therefore relates in the first place to an all-solid-state electrochemical generator comprising a negative electrode capable of delivering a lithium cation, an all-solid-state polymeric electrolyte formed from a macromolecular material in which an ionized lithium salt is dissolved and a positive electrode capable of incorporating the nonionized species corresponding to said lithium cation, characterized in that the all-solid-state polymeric electrolyte comprises at least one (where appropriate, several) fluoropolymer(s) in a macromolecular material/fluoropolymer(s) mass ratio of between 6 and 700.

In the current state of analysis of the experimentally observed phenomenon, it seems that fluorocompounds react according to an acid-base reaction by the substitution of the oxygen-containing species (oxide, hydroxide, carbonate) and/or of the nitrogen with fluorine. The fluorocompounds react in particular, according to this hypothesis, with lithium hydroxide and/or lithium oxide.





Document FR-A-2 563 382 describes various material formulations of the positive electrode based on  $V_2O_5$  and on metallic sulfide and oxide.

5 Preferably, the positive electrode will have a thickness of between 10 and 150  $\mu\text{m}$  and a proportion of active substance of between 20 and 80 wt%.

More specifically, the positive electrode will very preferably have a thickness of between 10 and 100  $\mu\text{m}$ , very advantageously between 20 and 100  $\mu\text{m}$ , and a proportion of active substance of between 25 and 65 wt%, very advantageously between 30 and 65%, or even between 45 and 65%.

15 To control the tree-forming phenomenon even more effectively, it has been found, unexpectedly, that it is advantageous for an antioxidant compound to be present in the polymeric electrolyte.

20 Although this amount of antioxidant may vary within appreciable proportions depending on the nature of the polymer used, it will be advantageous to use a proportion of antioxidant compound of between 0.5 and 25 3% with respect to the mass of polymer. It is obvious that this antioxidant must be compatible with said polymer.

Among antioxidants suitable within the context of the present invention, mention may be made of CHIMASSORB® 119, sold by Ciba-Geigy. Mention may also be made of quinone or hydroquinone derivatives and phenolic antioxidants.

35 Advantageously, the all-solid-state polymeric electrolyte includes a significant proportion of magnesia, between 5 and 30 wt%, preferably between 8 and 25 wt%.





Example 1

The single figure is a schematic sectional view of an electrochemical generator.

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The single figure shows the battery 1 consisting of a positive electrode 2, an electrolyte 3 and a lithium negative electrode 4, these three elements being produced according to the invention, and a collector 5 associated with the positive electrode, the lithium negative electrode 4 acting as collector.

15 The positive electrode is a composite electrode comprising a mixture of vanadium oxide, electrolyte and acetylene black, with an acetylene black content of 12 vol%.

20 The electrolyte consists of 69.7% of polyethylene oxide of 300 000 molecular mass, in which lithium trifluoro-sulfonylimide, in a proportion such that the oxygen/lithium atomic ratio is equal to about 20 (i.e. 17.6%), 9.8% of magnesium oxide, 0.7% of antioxidant (IRGANOX) and 2.2% of PVDF/HFP copolymer are dissolved.

25 The positive electrode has a capacity of about 1 mAh/cm<sup>2</sup> for a thickness of 60 µm.

30 The thickness of the negative electrode, the surface of which is very uniform, is 50 µm and the thickness of the electrolyte polymer is 50 µm.

Said electrochemical generator shows no significant variation in the specific energy after 300 charge/discharge cycles.

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Example 2

In this example, a comparison is made between the performance of two generators. Battery 2 is identical

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to battery 1 mentioned in example 1. Battery 3 consists of a positive electrode and a negative electrode identical to that of battery 2. The electrolyte of battery 3 consists of 71.3% of polyethylene oxide of  
 5 300 000 molecular mass, in which lithium trifluoro-sulfonylimide, in a proportion such that the oxygen/lithium atomic ratio is equal to about 20 (i.e. 18%), 10% of magnesium oxide and 0.7% of antioxidant (IRGANOX) are dissolved. Its thickness is 50  $\mu\text{m}$ .

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The sole difference between these two generators is the presence of PVDF/HFP copolymer, which is found in the electrolyte of battery 2.

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These two batteries are cycled with a controlled current density. The charge time is 10 hours and the discharge time is 5 hours. The current density is increased gradually until the maximum capacity of the battery is reached or until a short circuit due to the  
 20 formation of a dendrite occurs.

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In the case of battery 3, a short circuit appears when the charge current density exceeds  $0.1 \text{ mA/cm}^2$ . In the case of battery 2, it is possible to apply a charge current of  $0.2 \text{ mA/cm}^2$  without causing a short circuit. The maximum capacity of the battery is then reached.

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The use of the fluoropolymer therefore makes it possible to charge the battery at higher current densities.